## " VIJAYASREE" - PALAKKAD

First Mid Term Assessment - August 2018

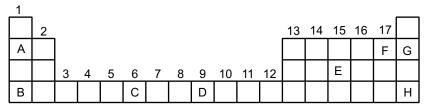
Std X	CHEMISTRY	Time 1 hr/ Score 20
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1.	<sub>21</sub> Sc: [Ar] 3d'4s <sup>2</sup>	
	<sub>12</sub> Mg :	(1)

- 2. The element Fe with atomic number 26 undergoes chemical reaction to form an ion with oxidation number +2 (4)
  - a) Write down the sub shell electronic configuration of this ion
  - **b)** Write down the chemical formula of one compound formed when this element react with chlorine (17 CI)
  - c) Can this element show variable valency? Why?
- **3.** Write the sub shell electronic configuration of the elements X, Y, Z and complete the table.

Element	Atomic No	Period	Group	Block
Х	12			
Υ	15			
Z	21			

4. (4)



- **a)** An incomplete form of periodic table is given . Which are the elements that belong to S block.?
- **b)** Which is the largest atom .?
- c) Which are the transition elements.?
- d) Which is the elements having highest electro negativity .?
- 5. Tabulate the following statements into s, p, d and f blocks

(4)

(6)

- a) Form coloured compounds
- b) Used as catalysts in petroleum industry
- c) Known as transition elements
- **d)** Include lanthanoids and actinoids
- e) Lose electrons in chemical reaction
- f) high ionization energy
- g) Low ionization energy
- h) Include elements from group 13 to 18
- Write the sub shell electronic configuration of elements P having atomic number 29.